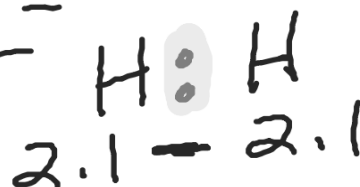


Non-Polar Covalent

Even sharing of e^-

H_2 $en = 2.1$

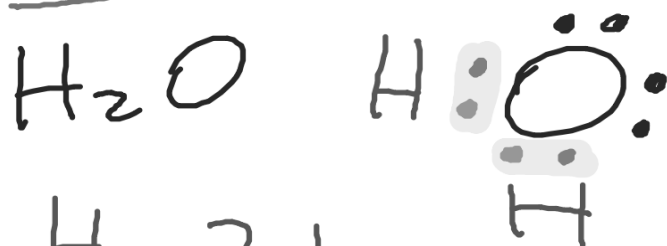


end electronegativity
difference

$$NP\ 2.1 - 2.1 = 0$$



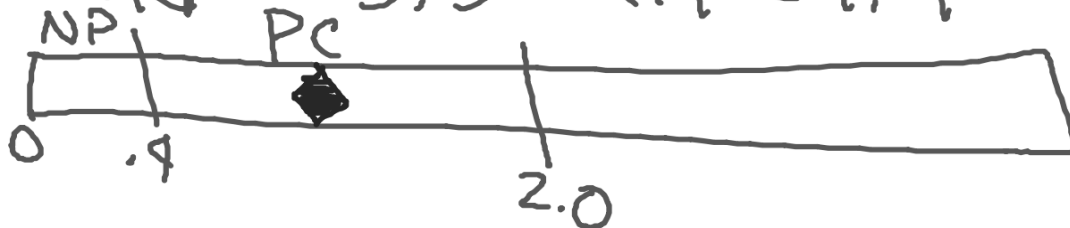
POLAR COVALENT
UNEVEN SHARING OF e^-

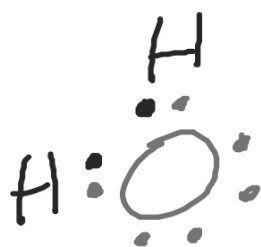


$$en H = 2.1$$

$$en O = 3.5$$

$$end = 3.5 - 2.1 = 1.4$$



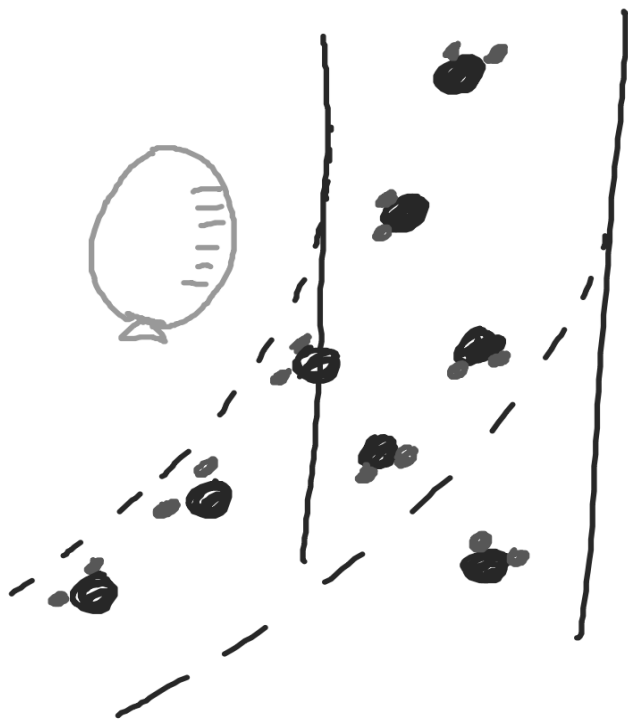


δ^+



δ^-

Partial
Difference



IONIC
STRONGER ATOM steals e^-

HF

$$en H = 2.1$$

$$en F = 4.0$$

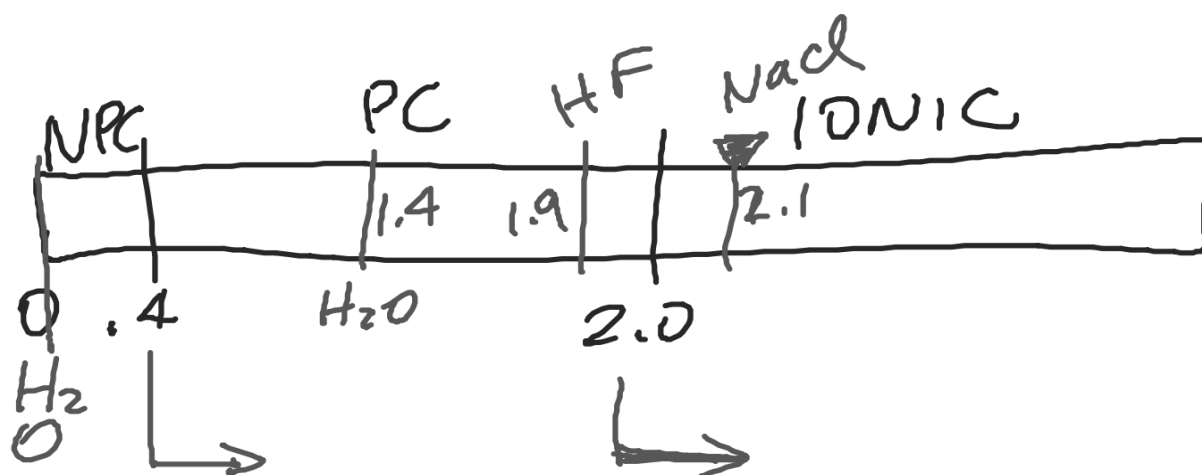
$$end = 4.0 - 2.1$$
$$1.9$$

NaCl

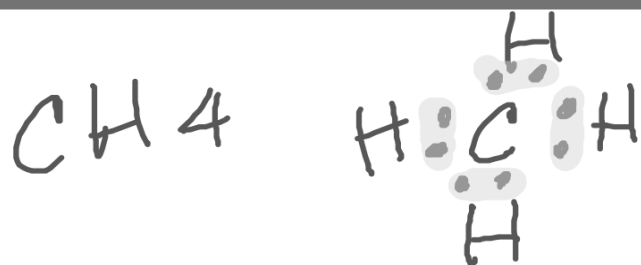
$$en Na = .9$$

$$en Cl = 3.0$$

$$end = 3.0 - .9 = 2.1$$



ON The lines Always goes
to higher zone



$$enC = 2.5$$

$$enH = 2.1$$

$$\begin{array}{r} 2.5 \\ - 2.1 \\ \hline .4 \end{array}$$

